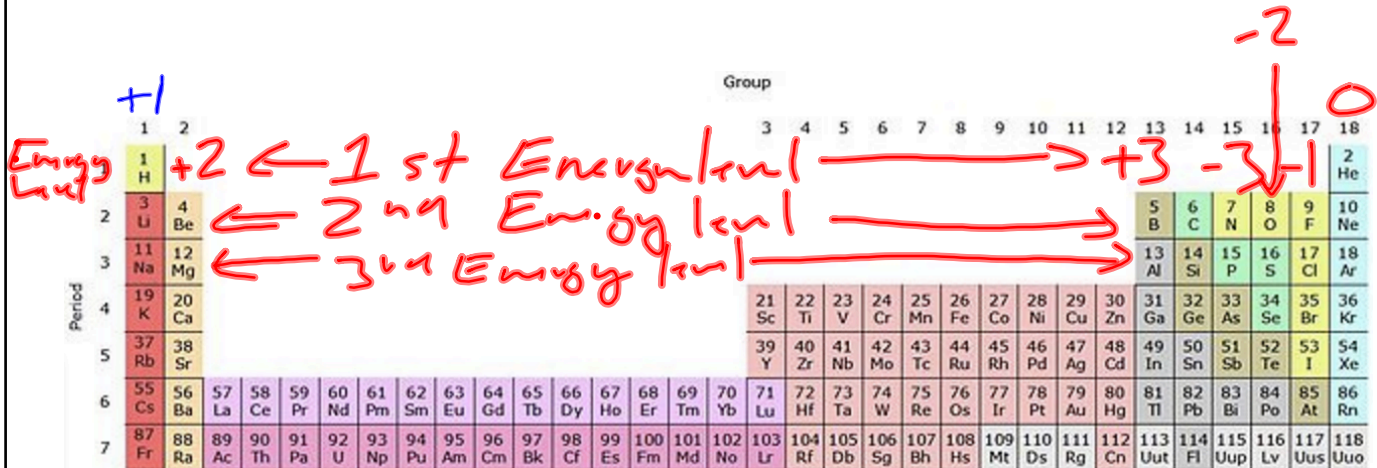


# Periodic Table Trends



1. organized by atomic number (number of protons)
2. Each row adds another energy level in order to hold the electrons in a balanced atom

## Electronegativity:

the ability of an atom to attract electrons

Scale: 0.7 (cesium) to 4.0 (Flourine)

prediction: higher on the right?

## Ionization Energy:

the energy required to remove an electron from an atom

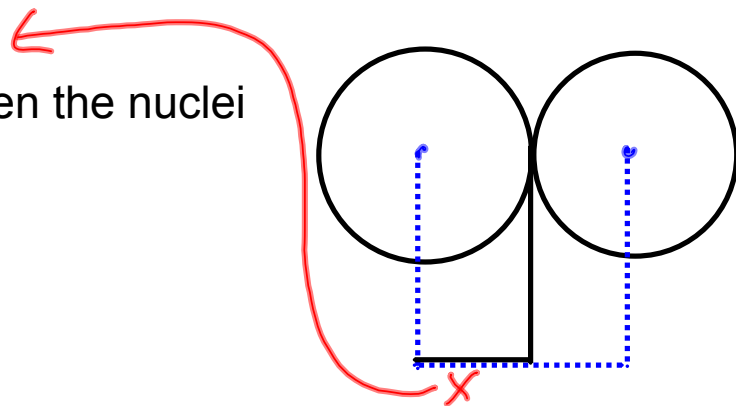
Prediction: Higher on the right side?

## Atomic Radius:

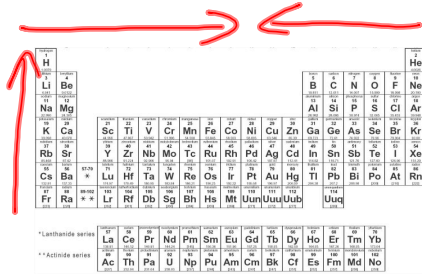
Half the distance between the nuclei

of two atoms

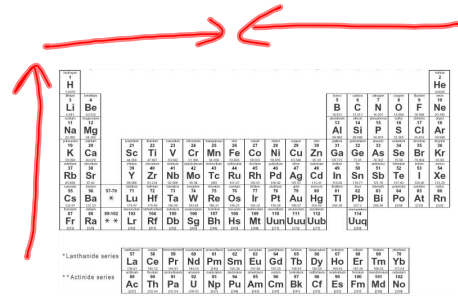
## Ionic Radius:



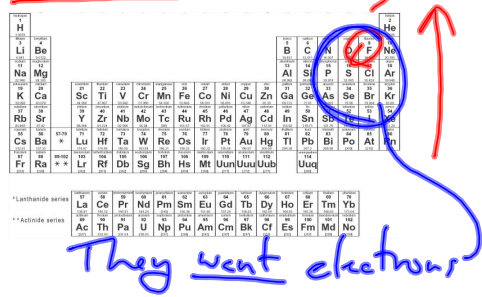
### Melting Point



### Boiling Point

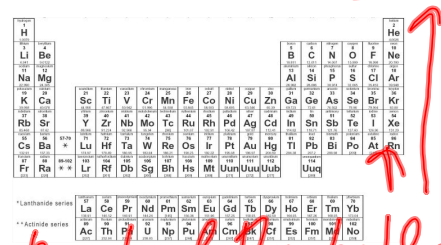


### Electronegativity



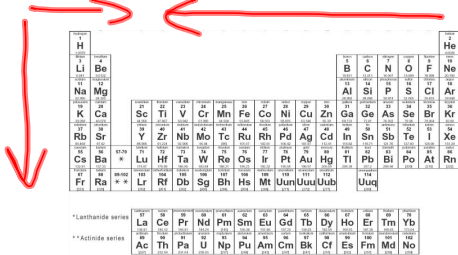
*They want electrons*

### Ionization Energy

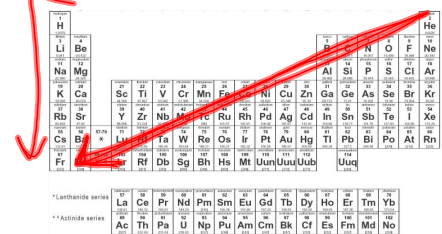


*how hard they hold onto the electrons they have*

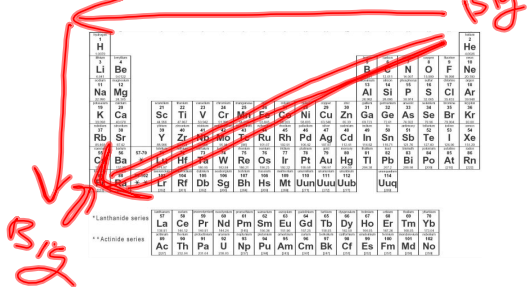
### Density



### Atomic Radius

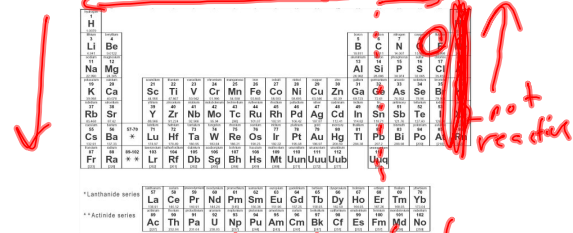


### Ionic Radius



*B<sub>15</sub>*

### Reactivity



*Think about electrons in the outer most energy level!*

### Metallic Character

