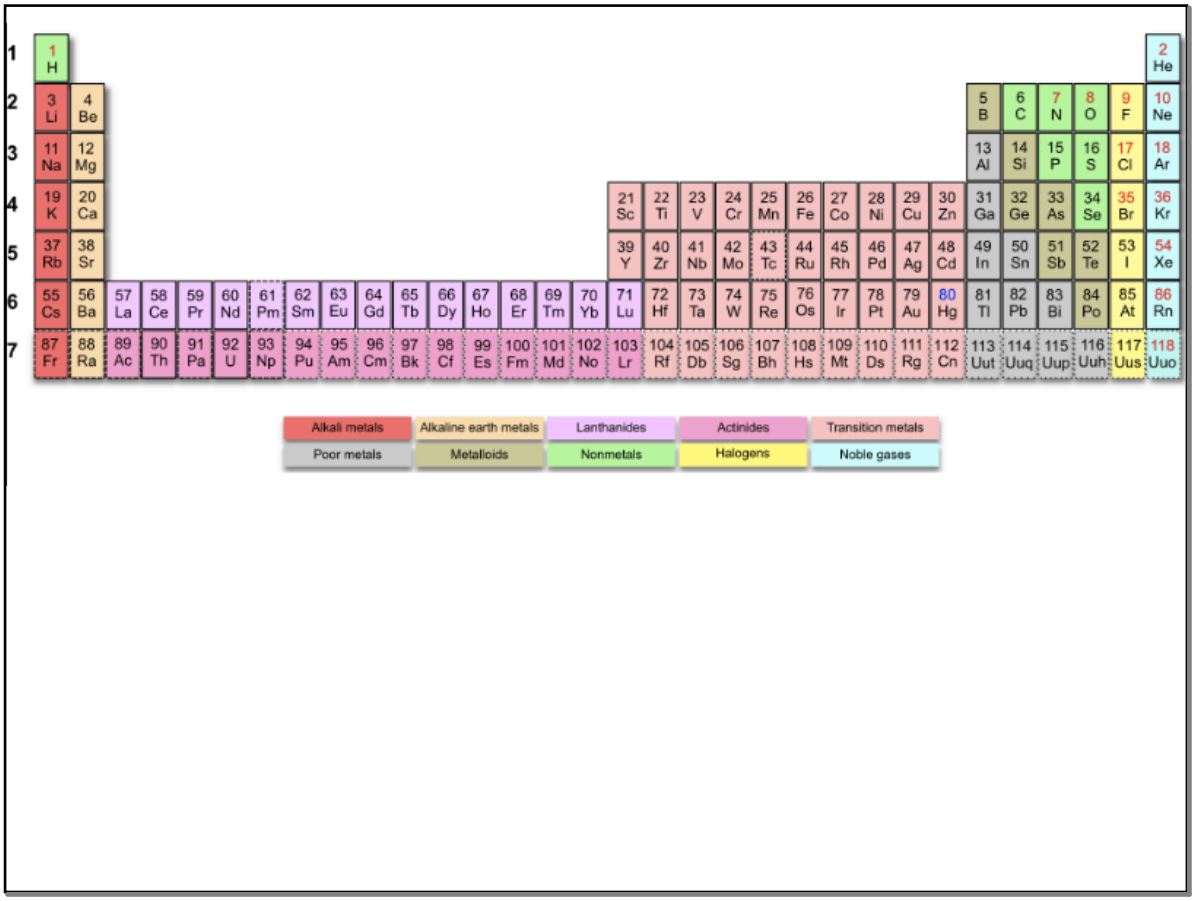
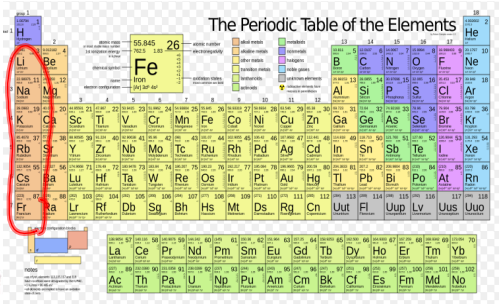


Alkali Metals



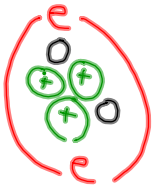
soft

low melting and boiling temperatures

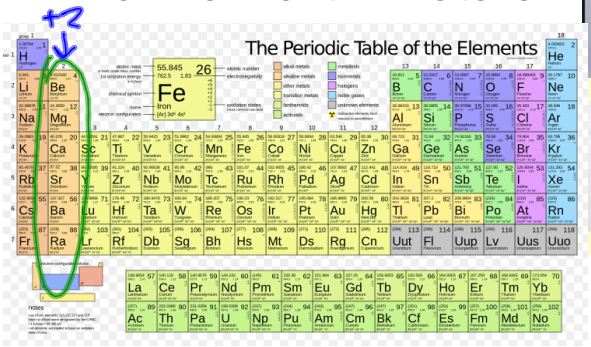
The MOST reactive

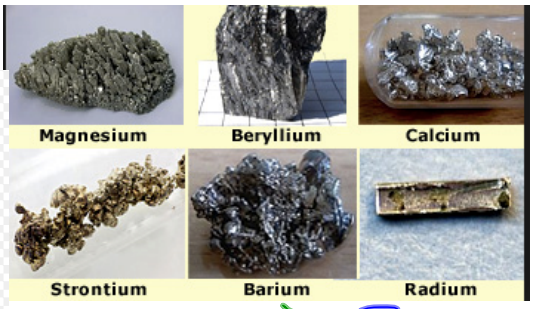
reactive in water

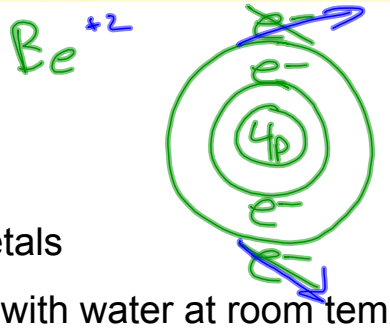
+1 Ion's because it drops the last electron



Alkaline Earth Metals







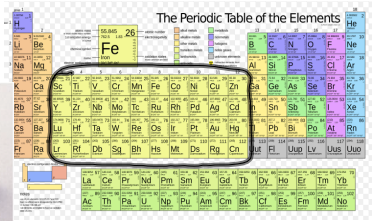
Physical: shiny and silvery

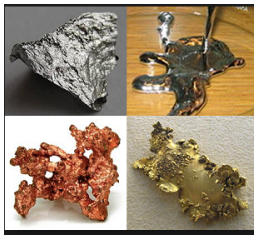
harder/more dense then alkali metals

reactivity: less then Alkali but reacts with water at room temp and boiling water

Ionization: +2 Cations

Transition Metals





Physical:

- Hard, high melting and boiling point
- High density
- Good conductors
- Ductile: (laffy taffy) Appear hard but bends without breaking

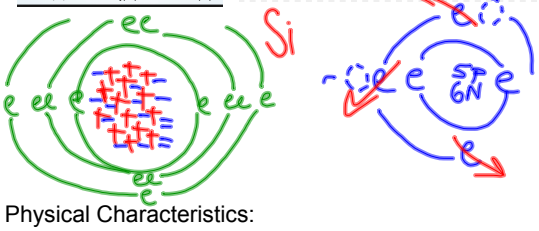
Reactivity: less reactive then alkali metals

- lower reactivity than alkaline metals and halogens
- higher reactivity then noble gasses

Ion Formation: lose electrons and become Cations

- "Super Sharers" Give away lots!!!
  - +1, +2, +3, +4

Metalloids



Physical Characteristics:

- Metallic
- Brittle
- Solid
- Semi-conductor: good conductors when they are hot

Reactivity: depends on what is reacting with

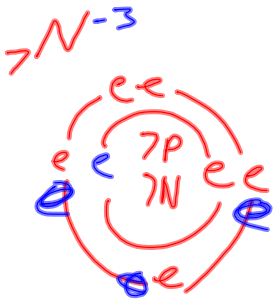
- not as reactive as other metals

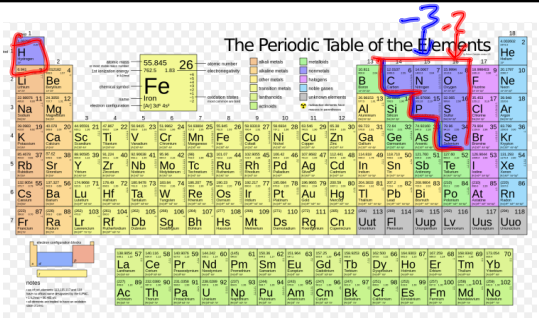
No Group Number: (13-16)

Ion formation depends on the column

- boron's column is +3, C's column is +/- 4...

Non-Metals





Group numbers 1, 14-16

poor conductor

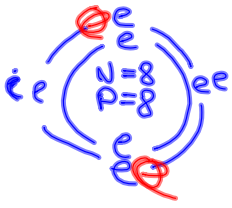
brittle

dull - does not reflect light

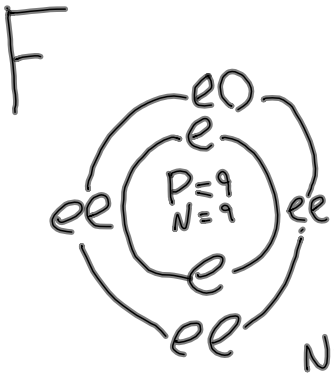
fewer energy levels = more reactive

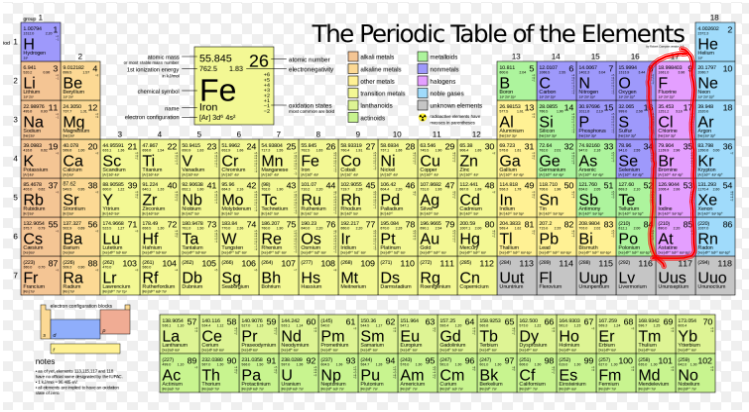
more energy levels = less reactive

Ions, +/-4, -3, -2 mostly Anions



Halogens



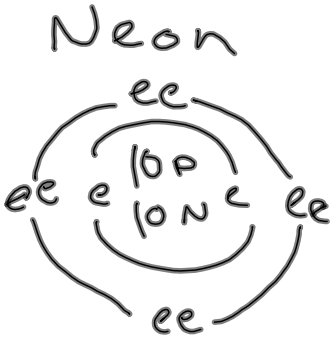


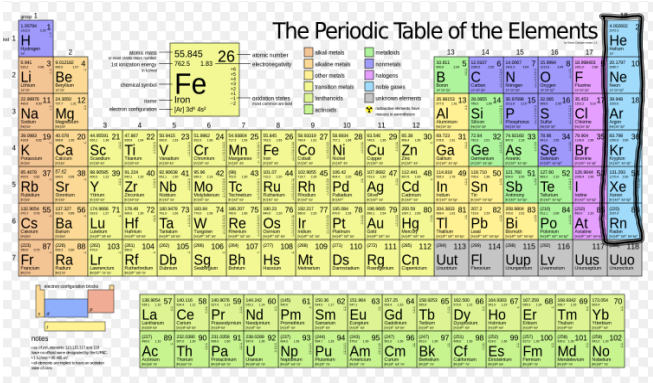
Non-metallic, toxic

Very reactive, less as you go down the column

Anions, gain 1 electron

Noble Gasses





Group: #18

Odorless, colorless gasses

Very Low reactivity

Don't take electrons from any one

Inner Transition Metals

Built in labs

Man made

Lots of different colors

tend to be radioactive!!!!!

Not reactive b/c man made

Random electron locations because

it doesn't matter