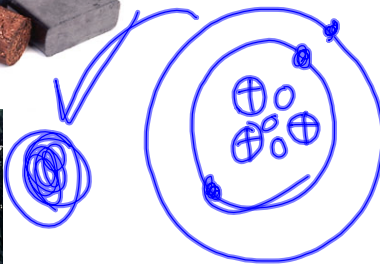
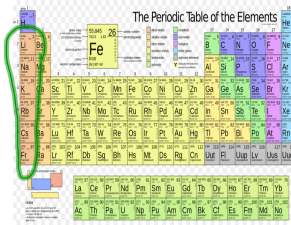


Alkali Metals

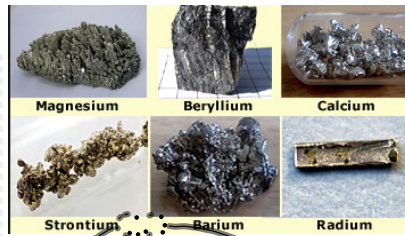


- Low melting and boiling points
- excellent conductivity of electricity
- Dissolves in water
- Reacts to Oxygen and nitrogen and nitro
- The most reactive metals

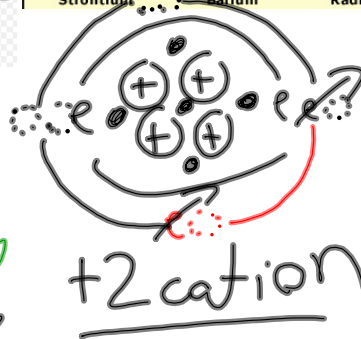


Alkaline Earth Metals

The Periodic Table of the Elements



PC:
 silver
 shiny
 low density
 very reactive
 has sens & water
 beryllium
 calcium
 strontium
 radium



Transition Metals



The Periodic Table of the Elements

Group # 3-12

Physical

Rock Based- high density

shiny

Malleable - Ductile - (laffy Taffy) Bend and not break

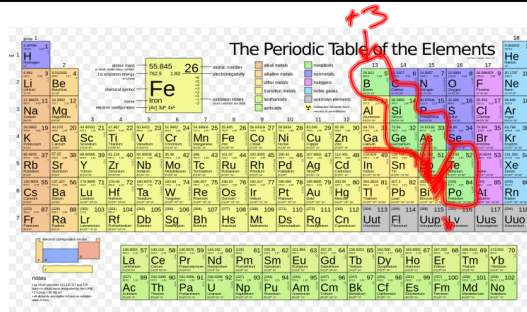
Conduct electricity

Reactivity: less than Alkali and Alkaline earth

Ion Formation: Cations, but not consistent (+1, +2, +3, +4
 Depending on their group)

Nickle, Iron and Gold

Metalloids



Group 13-16

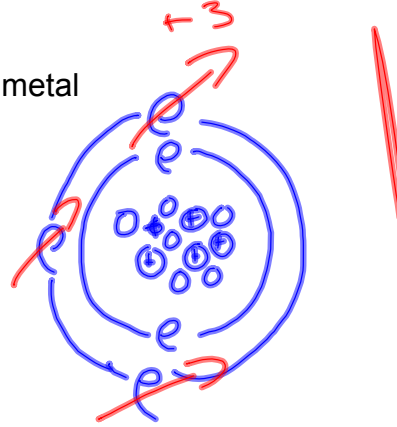
metallic

Conducts electricity when hot

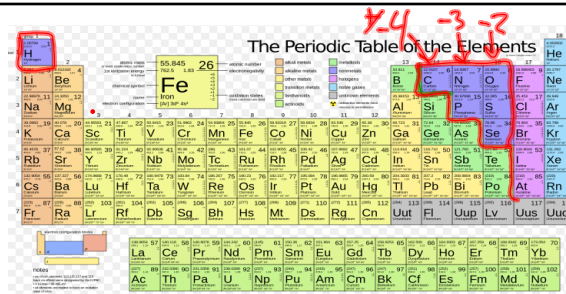
reactivity: acts like non-metal or a metal

Not as reactive as other metals

No single Ion Formation



Non-Metals



Group: 14-16

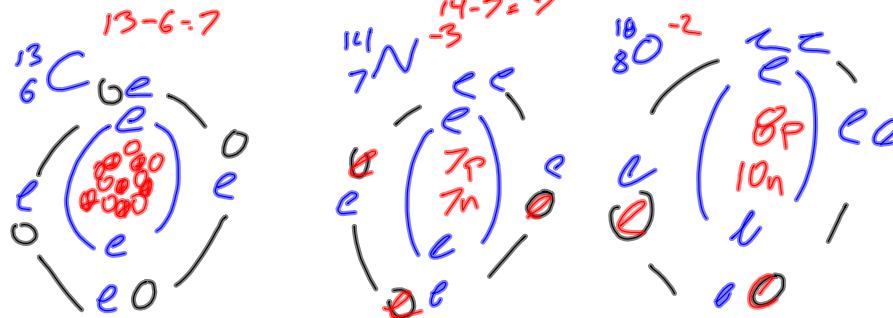
Not dense

do not conduct heat or electricity, found in organic stuff

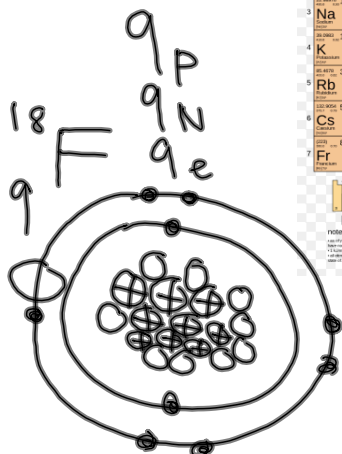
very brittle,

reactivity: decreases as you go down the table

Reacts by gaining electrons --> form Anions (-)



Halogens



Group # 17

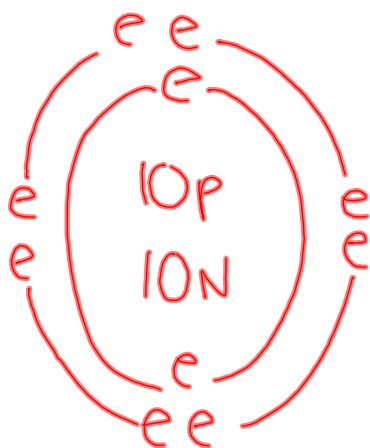
Non-metallic, solid liquid or gas
poisonous.

reactivity: THE MOST REACTIVE

never happy, they want more
electrons!!!!!!

ion formation: - 1 anions

Noble Gasses



odorless, colorless, non-metals

Group #18

Very Low reactivity

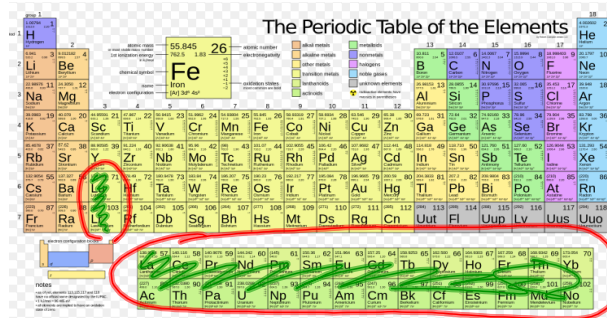
Inert - don't react

ion formation: none

Neon, Argon, Helium

Inner Transition Metals

Not Natural: they are made in the lab



Non-reactive but super radioactive

57: Lanthanum 2,8,18,
18,9,2

